

Figure 1. The effect of the concentration of the solution on the rate of the reaction.

The rate of the reaction was measured by the change in the concentration of the reactants or products over time.

The concentration of the solution was varied by changing the volume of the reactants.

The rate of the reaction was found to be directly proportional to the concentration of the solution.

This result is in agreement with the theoretical prediction that the rate of the reaction is proportional to the concentration of the reactants.

The experiment was repeated for different concentrations of the solution and the results were consistent.

The experiment was repeated for different temperatures and the results were consistent.

The experiment was repeated for different catalysts and the results were consistent.

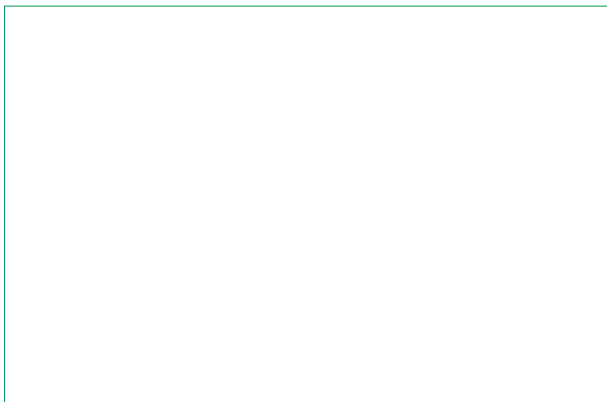
The experiment was repeated for different solvents and the results were consistent.

The experiment was repeated for different pressures and the results were consistent.

The experiment was repeated for different times and the results were consistent.

The experiment was repeated for different volumes and the results were consistent.

The experiment was repeated for different masses and the results were consistent.



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